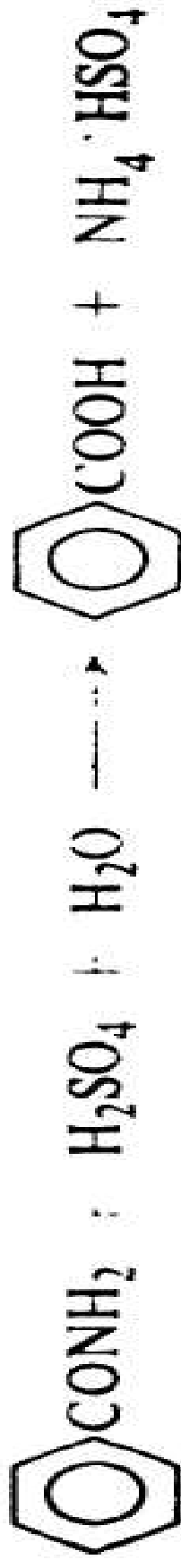


3-Formation of aldehydes by reduction.



2-Reaction of amide

1. Hydrolysis



Benzamide

Benzoic acid

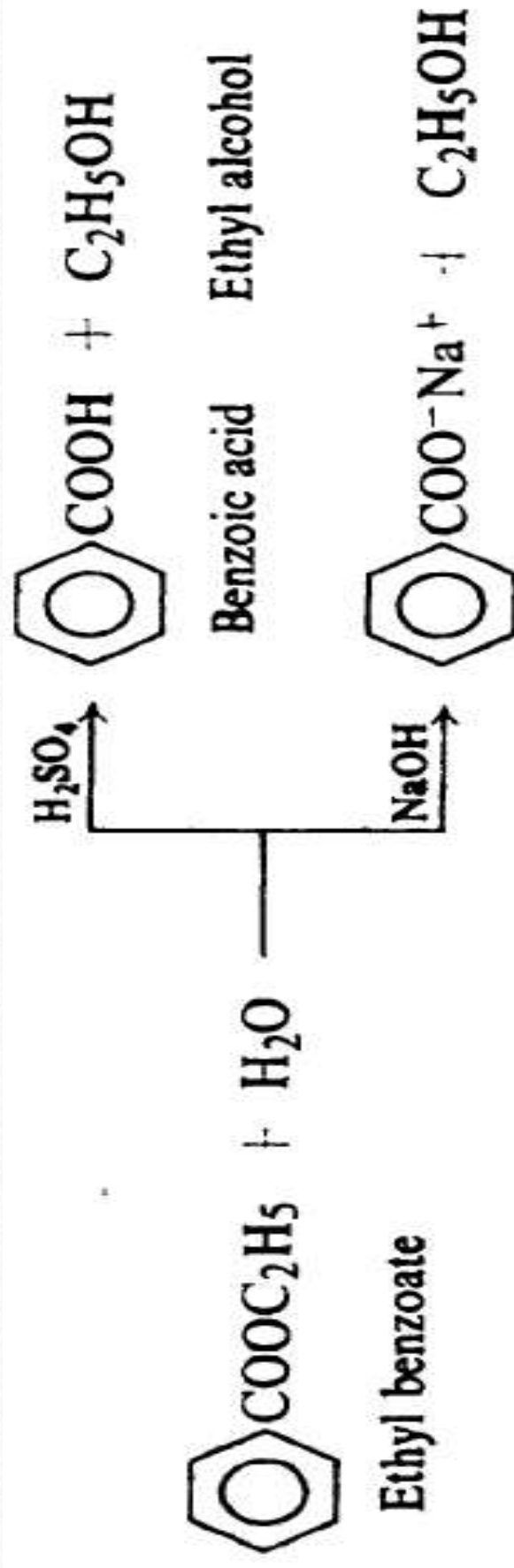


Butyramide

Sodium butyrate

2-Reaction of ester

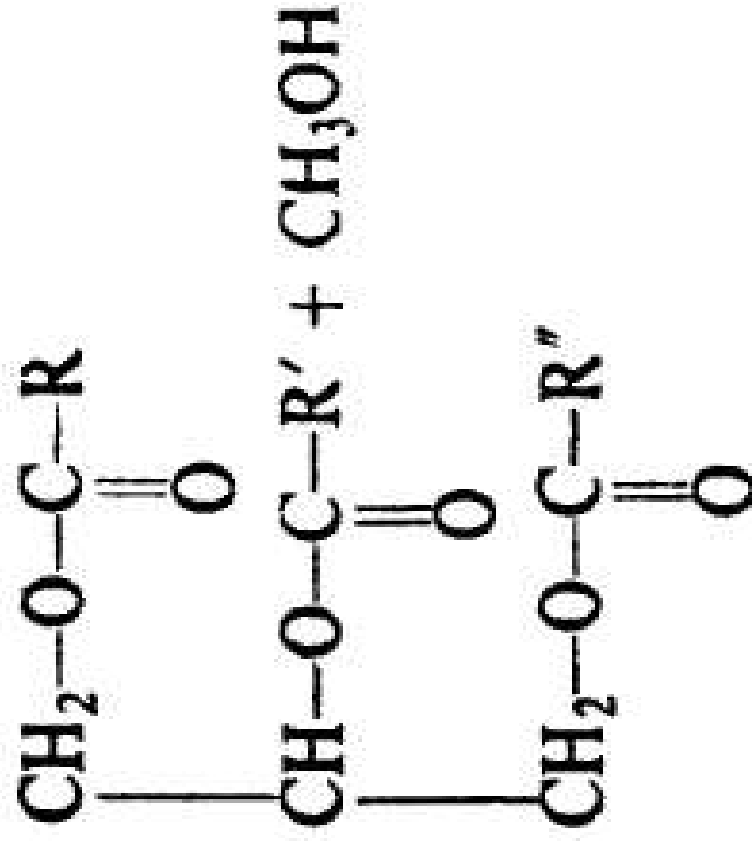
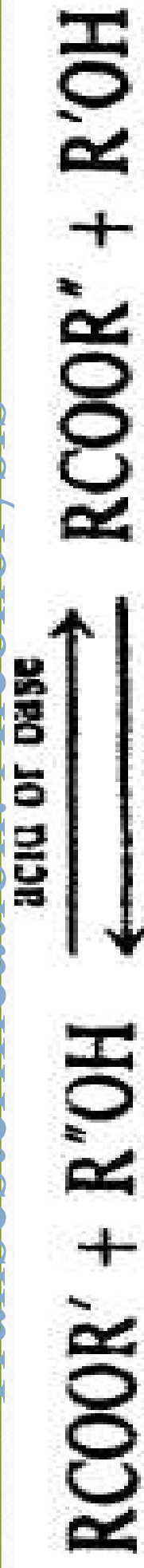
(a) Conversion into acids. Hydrolysis



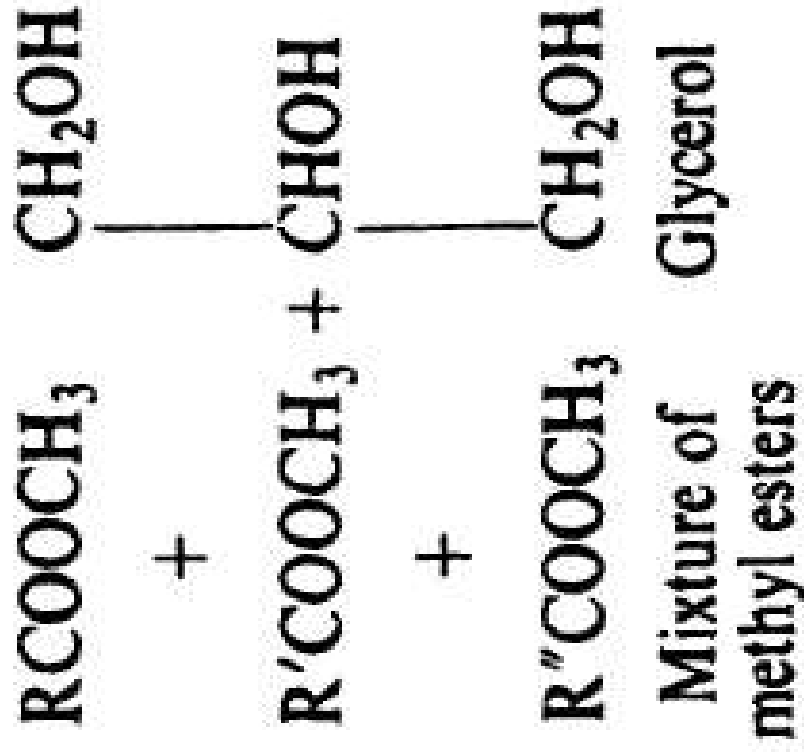
(b) Conversion into amides. Ammonolysis



Transesterification. Alcoholysis



A glyceride
(A fat)

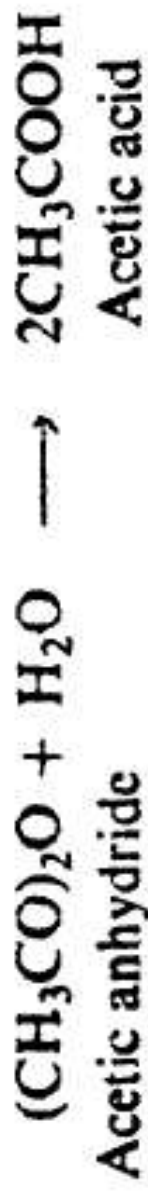


3-Reaction of acid anhydrides

1. Conversion into acids and acid derivatives



(a) Conversion into acids. Hydrolysis

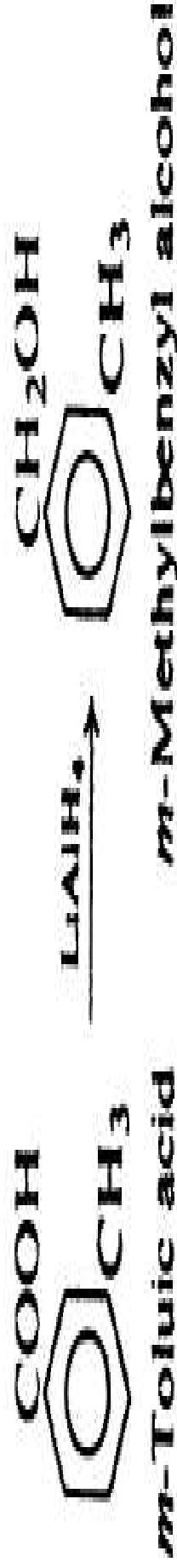


(b) Conversion into amides. Ammonolysis



3. Reduction

One of the few reducing agents capable of reducing an acid directly to an alcohol is lithium aluminum hydride, LiAlH_4

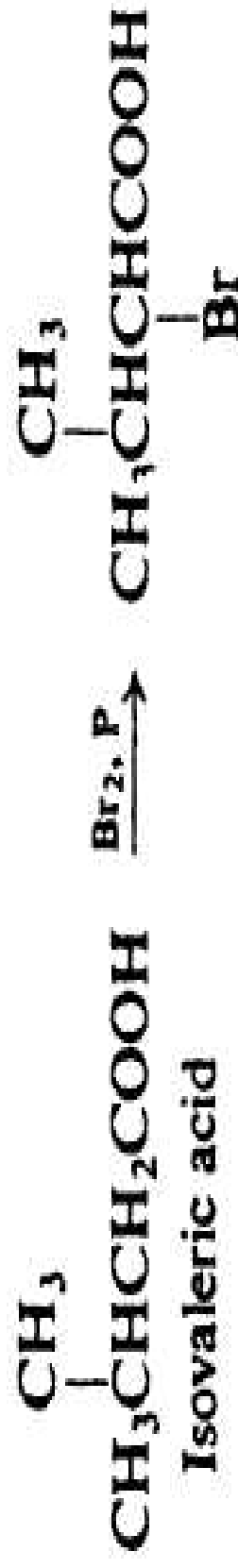


4. Substitution in alkyl or aryl group

α - Alpha-halogenation of aliphatic acids. Hell-Volhard-Zelinsky reaction



An α -haloacid



Dicarboxylic acids

If the substituent is a second carboxyl group, we have a dicarboxylic acid. For example:



Malonic acid

Succinic acid

Adipic acid

Propanedioic acid

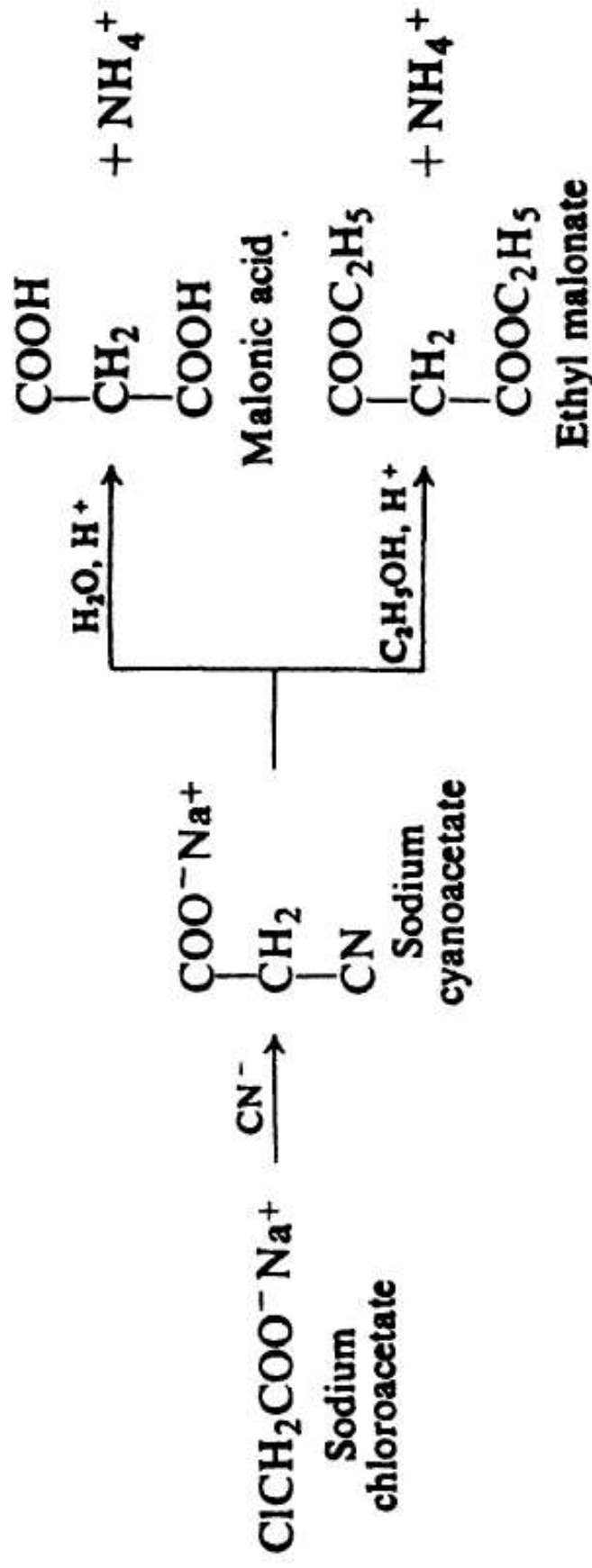
Butanedioic acid

Hexanedioic acid

- Table 2 Dicarboxylic acid

Name	Formula
Oxalic	$\text{HOOC}-\text{COOH}$
Malonic	$\text{HOOCCH}_2\text{COOH}$
Succinic	$\text{HOOC}(\text{CH}_2)_2\text{COOH}$
Glutaric	$\text{HOOC}(\text{CH}_2)_3\text{COOH}$
Adipic	$\text{HOOC}(\text{CH}_2)_4\text{COOH}$
Maleic	<i>cis</i> - $\text{HOOCCH}=\text{CHCOOH}$
Fumaric	<i>trans</i> - $\text{HOOCCH}=\text{CHCOOH}$

Preparation of Dicarboxylic acid



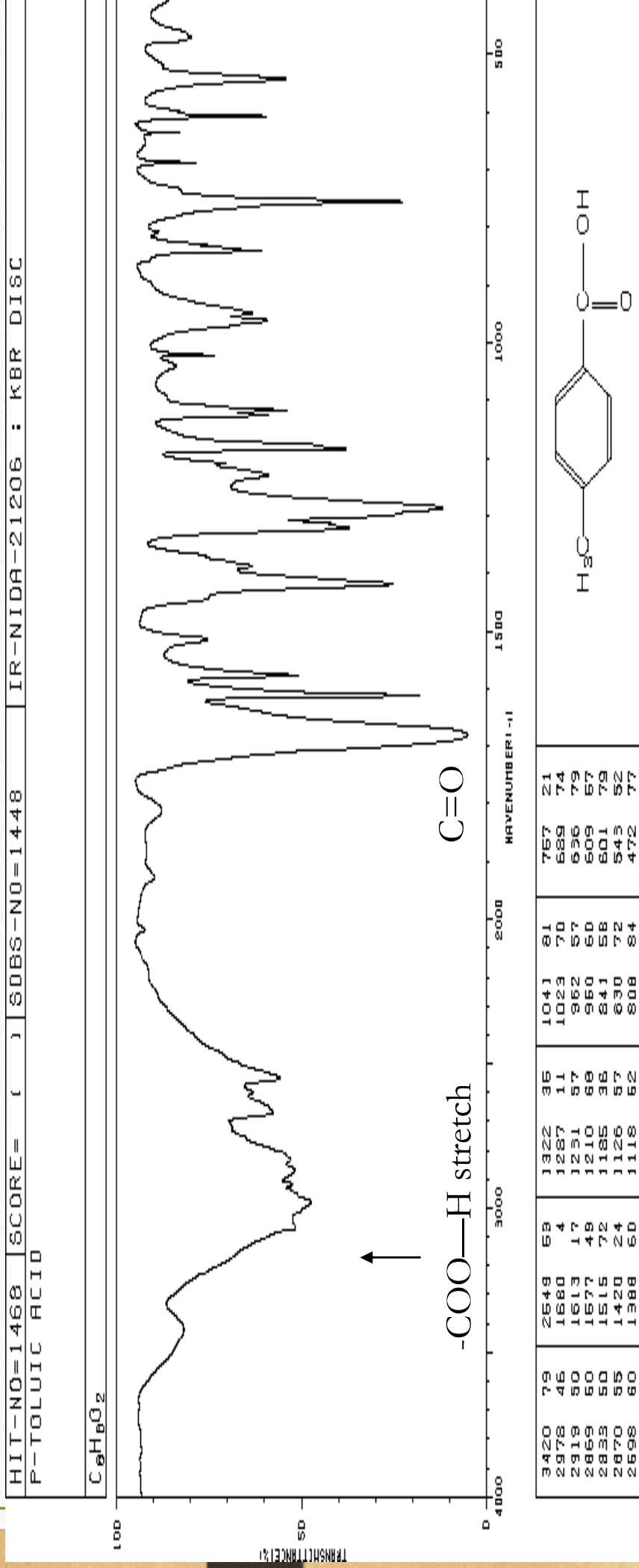
spectroscopy

- IR: -COOH O-H stretch $2500 - 3000 \text{ cm}^{-1}$ (b)

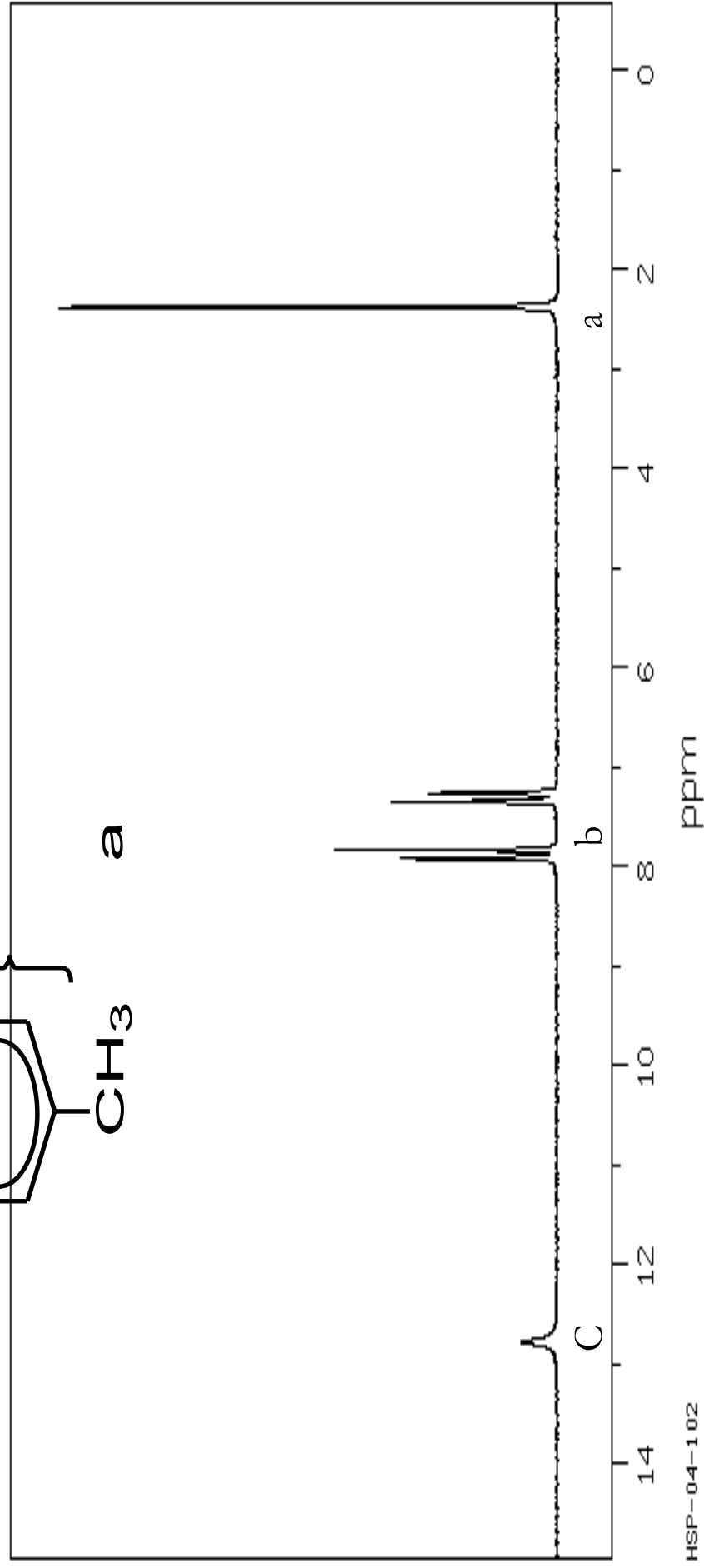
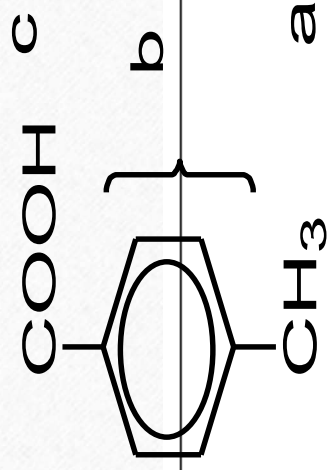
C=O stretch $1680 - 1725$ (s)

- nmr: -COOH $10.5 - 12$ ppm

IR SPECTRUM of *p*-toluic acid



NMR Spectrum of p-toluic acid



HSP-04-102