

## ELECTRICAL AND OPTICAL PROPERTIES OF CHEMICALLY DEPOSITED $\text{SnO}_2$ :I COATINGS

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Received 24 October 1987; in revised form 9 May 1988

Thin coatings of  $\text{SnO}_2$  and  $\text{SnO}_2$ :I are prepared using the spray pyrolysis technique. The effect of iodine doping on the electrical and optical properties of  $\text{SnO}_2$  are studied and compared with other types of doping. The electrical sheet resistance is minimum ( $30 \Omega/\square$ ) and optical transmission at solar maximum is 95% for an  $\text{SnO}_2$  coating doped with 0.3 wt% iodine. The absorption coefficient data versus photon energy for  $\text{SnO}_2$  coatings doped with 0.3 wt% iodine are analyzed and interpreted in terms of direct and indirect allowed transitions.

### 1. Introduction

A transparent-conductive semiconductor possesses the unique properties of good electrical conductivity and high optical transparency in the visible spectrum [1-8]. The optical and electrical properties of  $\text{SnO}_2$  coatings are improved by doping with fluorine, chlorine, and antimony [9-14,7]. The investigations of  $\text{SnO}_2$  coatings are interesting because of their wide range of applications: (i) as a heat mirror suitable for application in solar photo-thermal conversion [6], (ii) heterojunctions of transparent-conductive oxides on bulk semiconductors show considerable promise for use in inexpensive solar cells [2,15], (iii) the refractive index of  $\text{SnO}_2$  is in the right range so as to form an antireflection coating on silicon and GaAs [4]. In the present work  $\text{SnO}_2$  coatings doped with iodine were prepared using the spray pyrolysis technique. The electrical and optical properties were studied with various iodine doping concentrations.

### 2. Experimental

The coatings were chemically deposited using the spray pyrolysis technique onto heated borosilicate glass substrates. Spraying solutions containing 30 g  $\text{SnCl}_4$ , 37 ml  $\text{H}_2\text{O}$ , 30 ml  $\text{CH}_3\text{CH}_2\text{OH}$  and  $\text{HCl}$  were used for preparing  $\text{SnO}_2$  coatings.  $\text{HCl}$  was